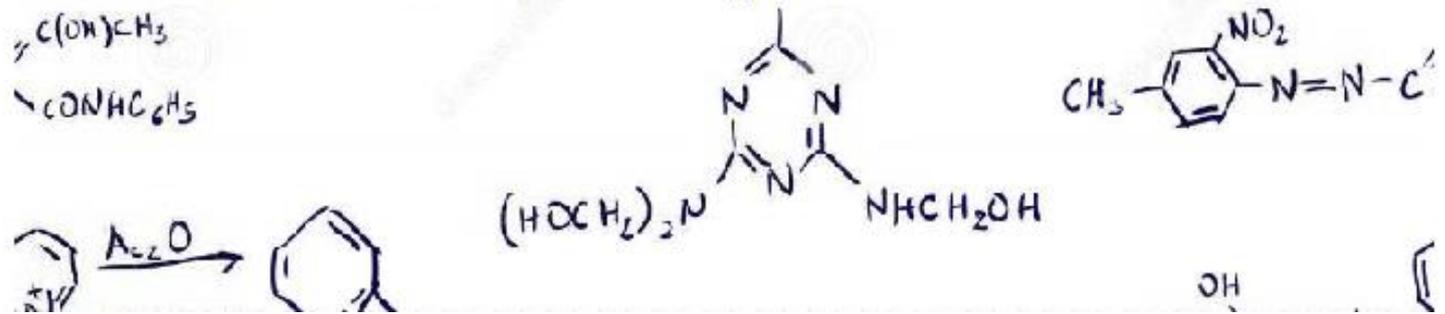
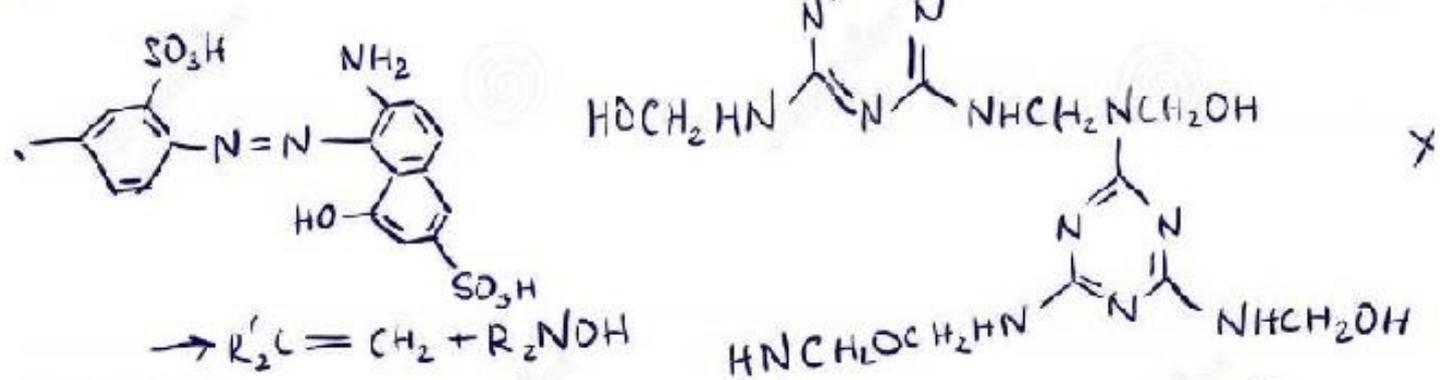
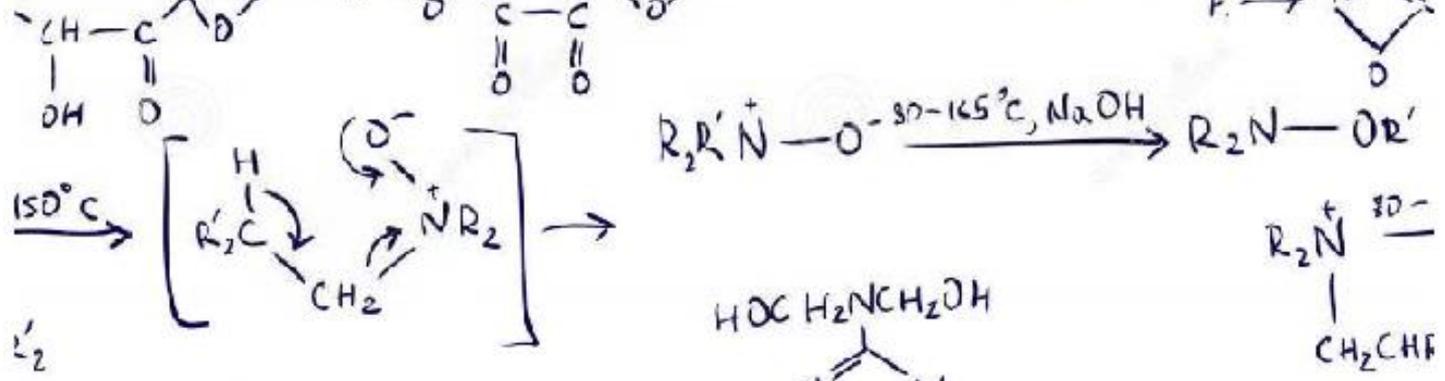
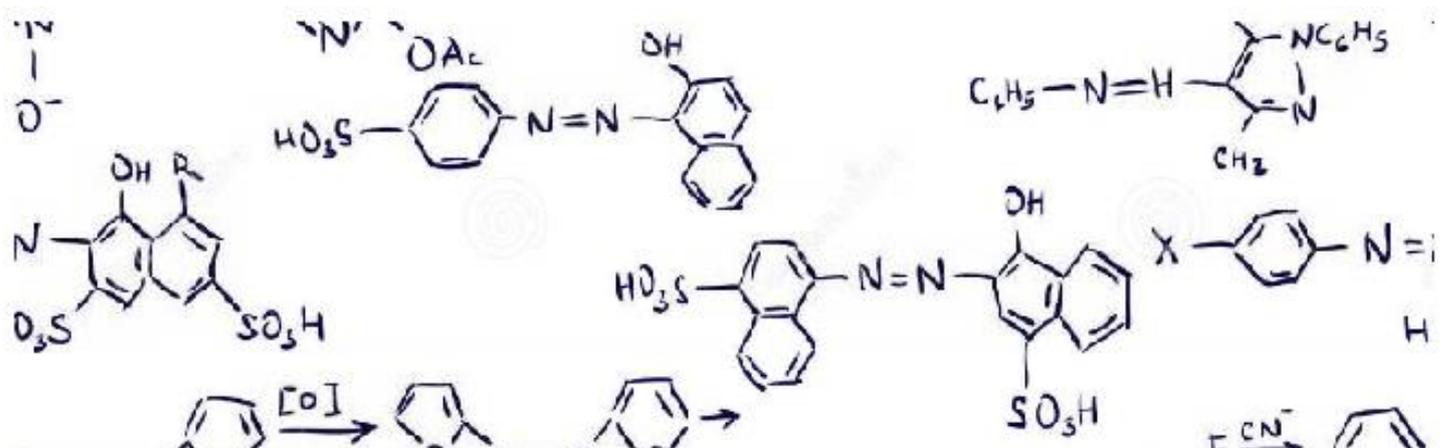
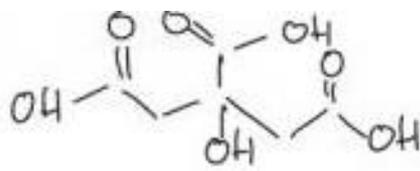
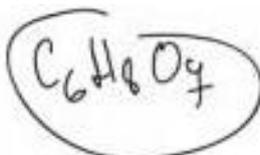


$$\begin{aligned}
 & \mathbb{P}\{N \leq \frac{1}{2} \mid X > A\} = \sum_{k=0}^{\infty} e^{-\frac{k^2 \pi^2}{2}} = H(k) \\
 & \int dG_k(x) \geq \frac{1}{2} \sum_{k=0}^{\infty} e^{-\frac{k^2 \pi^2}{2}} = H(k) \\
 & f_{n-1}(t) = \int_0^1 f_n(u) f_1(t-u) du = \frac{2^{n+1} t^n e^{-2t}}{n!} \quad \lim_{t \rightarrow 0} (f_n(t)) = 0 \\
 & \log \varphi(t) = i\gamma t - c |t|^\alpha [1 + i\beta \frac{t}{|t|} \omega(t, \alpha)] \quad B(u) = \sum_{k=1}^r \Psi^*(b_k u) \quad C_{iv} = \sum_{j=1}^n a_{ij} b_j v \\
 & \int_{-\infty}^{\infty} e^{-\frac{u^2}{2}} du = F(x) \left(\frac{1}{\sqrt{2\pi}}\right)^{-1} \quad |\Psi_S(t)| = \left| \int_{-\infty}^{\infty} e^{itx} dF(x) \right| \leq \int_{-\infty}^{\infty} e^{-\nu x} dF(x) = \varphi_S(i\nu) \\
 & \prod_m = \prod_r \prod_{m-r} \\
 & |X \cup Y| = |X| + |Y| - |X \cap Y| \quad \lim_{n \rightarrow \infty} \frac{1}{n} k_n \left(\frac{x}{\sqrt{n}}\right) = \frac{1}{\sqrt{2\pi}} e^{-\frac{x^2}{2}} \\
 & f: X \rightarrow X \cap W \\
 & \left( \sum_{k=1}^r P_k^\alpha \log_2 \frac{1}{P_k} \quad \left| \sum_{k=1}^r P_k^\alpha \log_2 \frac{1}{P_k} \right|^2 \right) \quad f g(u_i) = f \left( \sum_{j=1}^{\dim V} a_{ji} v_j \right) = \sum_{j=1}^{\dim k} a_{ji} \left( \sum_{k=1}^{\dim V} b_{kj} w_k \right) \frac{\binom{2k}{k}}{2^k} \approx \frac{1}{\sqrt{\pi k}}
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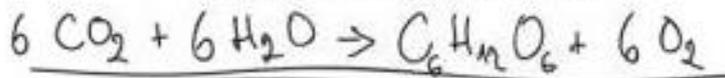
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192,15 g/mol

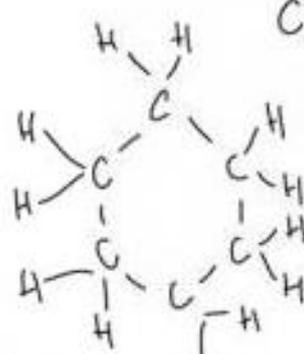
PHOTO SYNTHESIS



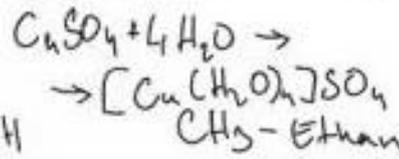
$H_2O_2$

$O_2$

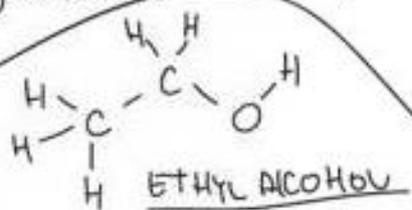
HCl  
NaCl



Cykllohexan



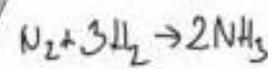
CH<sub>4</sub> - Methan



HNO<sub>3</sub> - Nitric acid

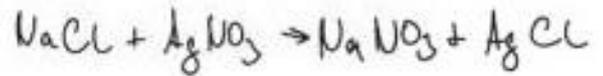
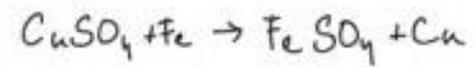
H<sub>3</sub>PO<sub>4</sub> - Phosphoric acid

H<sub>2</sub>SO<sub>4</sub> - Sulfuric acid

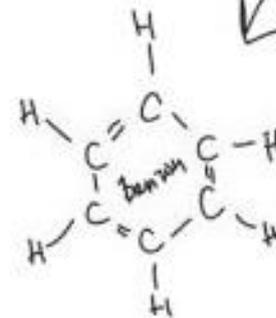


H<sub>4</sub>N<sub>2</sub>O<sub>3</sub>

CH<sub>2</sub>O - Glukosa



C<sub>6</sub>H<sub>6</sub> / CH - Benzol



$R-C(=O)NO_2$  c1ccc(cc1)/N=N/c2ccc(cc2)N(C)C  $I_2 + OH^- \rightarrow IO_3^- + I^- + H_2O$   $\frac{[O_2^{2+}] \cdot [S^{2-}]}{[CH_3S]}$   $= \frac{10^{-10} \cdot 10^{-4}}{10^{-5}} = 10^{-9}$

$C_6H_5CH_3 \xrightarrow{+3Cl_2, -H_2O} C_6H_5COOH$   $CH_3-C(=O)OH$   $CH_3-C(=O)OH + H_2O \rightleftharpoons CH_3-C(=O)O^- + H^+$

$WO_3 + 3H_2 \rightarrow W + 3H_2O$   $R-C(=O)OH + H_2 \rightleftharpoons R-C(=O)O^- + H^+$

$Ca(OH)_2 + Ca(HCO_3)_2 \rightarrow 2CaCO_3 \downarrow + 2H_2O$   $[Cu(NH_3)_4]^{2+} + H_2 \rightleftharpoons 2H^+ + 4NH_3$   $\Delta H = -455 kJ/mol$

$HOC-COOH \rightarrow CO + CO_2$   $R-COO^- + Ca^{2+} \rightarrow R-COO-Ca^+$

$Ca^{2+} + CO_3^{2-} \rightarrow CaCO_3$   $CH_3-C(=O)OH + H_2O \rightleftharpoons CH_3-C(=O)O^- + H^+$

$4HClO_2 \rightarrow 3HClO + HCl$   $2I^- \rightarrow I_2 + 2e^-$   $PbS + 4HCl \rightarrow 2H_2S + PbCl_2$

$n \cdot R \cdot T = p \cdot V$   $HOC-COOH \rightarrow HCl + CO_2$   $Fe_2O_3 + 3H_2 \rightarrow 2Fe + 3H_2O$   $Na_2Si_2O_7 + Ca^{2+} \rightleftharpoons CaSi_2O_7 + 2H^+$

$2NaBr + H_2SO_4 \rightarrow Na_2SO_4 + 2HBr$   $[Fe(CN)_6]^{3-}$

$2HNO_2 + 2HCl \rightarrow 2HNO + 2Cl_2 + 2H_2O$   $c(\frac{1}{2} HNO_2) = \dots$

$Cu \xrightarrow{HNO_3} Cu(NO_3)_2$   $T_{Cu} = \frac{m_{Cu}}{m} = \frac{0,635}{4} = 0,15875$

$\frac{1}{x} = R \cdot (2-n) \cdot (\frac{1}{r_1} - \frac{1}{n_1})$

$Cu + \frac{1}{2} O_2 \rightarrow CuO$

$2HNO_2 + 2HCl \rightarrow 2HNO + 2Cl_2 + 2H_2O$